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base. the particle that remains when an acid
has donated a hydrogen ion; OH^- is the
conjugate base of the acid water. Conjugate
acid-base pair. two substances that are
related by the loss or Page 1/22

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formed when a base gains a hydrogen ion:
 NH_4^+ is the conjugate acid of the base
 NH_3 *20.3. two substances that are related

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by the loss or gain of a single hydrogen ion.

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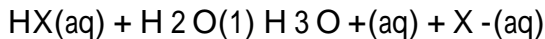
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Great Wall of China. Chapter 20 Acids Bases - modapktown.com Title: Chapter 20: Acids and Bases 1 Chapter 20 Acids and Bases 2 Acids. Give food a sour flavor ; Aqueous solutions of acids are electrolytes ; React with compounds containing hydroxide ions to form salt and water; 3 Bases. Feel slippery ; Taste bitter ; Are electrolytes ; React with acids Page 7/13

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Chapter 20: Acids, Bases, and Redox. 20.1 Acid and bases; 20.2 Relative Strengths of

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Acids and Bases; 20.3 Acidic, Basic, and Neutral Solutions; 20.A Acid Rain and Basic Oceans; 20.4 Losing and Gaining Electrons; 20.5 Electrochemistry; 20.6 Corrosion and Combustion

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Acid Base Conjugate Conjugate acid base
Conjugate acids and bases The forward reaction is the reaction of an acid and a base. The reverse reaction is also the reaction of an acid and a base.

Chapter 18: Acids and Bases

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CHAPTER 20. The Properties of Acids and Bases A base, according to the Bronsted-Lowry theory, ammonia acts as a base when it accepts hydrogen ions from water
Conjugate acid the particle formed when a base gains a hydrogen ion: NH_4^+ is the conjugate acid of the base NH_3
20.3 Chem 1120 - Chapter 15: Acids and Bases

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14.4 Name each of the following acids or bases: a. $\text{Al}(\text{OH})_3$ b. HBr c. H_2SO_4 d.

KOH e. HNO_2 f. HClO_2 14.5 Write formulas for each of the following acids and bases: a. rubidium hydroxide b.

hydrofluoric acid c. phosphoric acid d. lithium hydroxide e. ammonium hydroxide f. periodic acid 14.6 Write formulas for each

of the following acids and bases: a. barium hydroxide b.

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